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ChemActivity 6 1.The number of the valence shell corresponds to

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the number of the row  
of the Periodic Table.

2. Because the nuclei  
gain one proton as one  
moves from left to  
right across a period  
but the number of  
inner shell electrons  
remains constant.

### **Chem;GI;Answers - d bhs.wvusd.k12.ca.us**

6. IE of Kr > IE of Br  
because they are in the  
same valence shell and  
Kr has the higher core  
charge (+8 vs. +7). IE



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of Rb is the lowest because core charge is +1 and its valence shell ( $n = 5$ ) is larger than the valence shell ( $n = 4$ ) of Kr and Br.

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6. What is the significance of the atomic number,  $Z$ , above each atomic symbol in the periodic table? 7. Based on your answer to CTQ 6, what do all nickel (Ni) atoms have in common?

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ChemActivity 1 The  
Nuclear Atom 3

## Answers **The Nuclear Atom**

Answers to  
ChemActivity 1 - The  
Nuclear Atom- Chapter  
5 1. 6 2. 6,7,7 3. 6,6,7  
4. (a) neutral same e  
and p, ion different e  
and p (b) assign +1 to  
each proton and -1 to  
each electron and take  
the difference.

## **Chemactivity 7** **Answers**

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Periodic trends predict differences between elemental characteristics as you move across the periodic table. Trends are based on

Coulomb's law which mathematically relates several characteristics of an elements. Atomic size measured the distance between the nucleus of an atom and the outermost non-valence electrons of the atom. Atomic size

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decreases from left to right, because the addition of ...

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Group trends in atomic size (6.3) As the atomic number increases within a group the charge on the nucleus increases and the number of occupied energy levels increases. Explain the group trend of atomic size (6.3) Atomic size

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Answers

increases because the shielding effect is greater than the effect of the increase in nuclear charge.

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Atomic Structure. 1  
The Nuclear Atom 2. 2  
Atomic Number and  
Atomic Mass 8. 3  
Coulombic Potential  
Energy 16. 4 The Shell  
Model (I) 22. 5 The  
Shell Model (II) 30. 6

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Atomic Size 40. 7

Electromagnetic  
Radiation 44. 8

Photoelectron

Spectroscopy 48. 9 The  
Shell Model (III) 56. 10

Electron Configurations

60. 11 Electron

Configurations and the  
Periodic ...

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Atomic Radii Size: 1.

Explain atomic radius:

A measure of the size

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(volume) of an atom.

Determined by

measuring half of the distance between two identical atoms bonded together. 2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right. 3.

## **ANSWER KEY Trends on the Periodic Table**

*Page 16/26*



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Atoms, atomic number and size are factors that define an element and depending on the atomic size or number, an element can be placed as a solid, gas or liquid in a periodic table. This is a good challenge for you if you just studied the periodic table. Give it a shot!

### **Lesson 10 Test - Atomic Size And The Periodic Table ...**

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ChemActivity 2 Atomic  
Number and Atomic  
Mass Are All of an  
Element's Atoms

Identical?) WARM-UP

Model 1: The Average

Mass of a Marble In a

collection of four

marbles, 25% of the

marbles have a mass

of 5.00 g and 75% of

the marbles have a

mass of 7.00 g. 5.00 g

7.00 g 26.00 g The

average mass of a

marble can be

determined by dividing

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the total mass of the  
marbles by the total  
number of marbles ...

## **Solved: ChemActivity 2 Atomic Number And Atomic Mass Are A**

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Take Krypton for  
example. On top it is  
written 36 and down  
bottom it is written

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83.80. 36 is the atomic number and 83.80 is atomic weight or atomic mass. the easy way to remember is that the number on TOP is atomic number and bottom is the weight or mass. HOPE IT HELPED :)

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To the Student 1  
Atomic Structure 1 The  
Nuclear Atom 2 2

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Energy 16 4 The Shell

Model (I) 22 5 The Shell

Model (II) 30 6 Atomic

Size 40 7

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Shell Model (III) 56 10

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